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(unbalanced) equation determine the number of moles of $\text{NH}_3(\text{g})$ that can be produced from the following: a. 0.20 mol $\text{N}_2(\text{g})$ reacts completely with $\text{H}_2(\text{g})$. b. 0.30 mol $\text{H}_2(\text{g})$ reacts completely with $\text{N}_2(\text{g})$. 2. Considering the reaction represented by the (unbalanced) equation

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