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Chapter 12 Stoichiometry Practice Problems Chapter 12 Stoichiometry Practice Problems Answer Key A In any stoichiometry problem, the first step is always to calculate the number of moles of each reactant present. In this case, we are given the mass of K 2 Cr 2 O 7 in 1 mL of solution, which we can use to calculate the number of moles of K 2 Cr ...

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Answer Key Chapter 12: Stoichiometry Mole Ratios Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows: 4Al + 3O 2 → 2Al 2O 3 a. If you use 2.3 moles of Al, how many moles of Al 2O 3 can you make? b. If you want 3.9 moles of Al 2O 3, how many moles of O 2 are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate: 2Fe + 3H 2SO 4 → Fe 2(SO 4 ...

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